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# Room Temperature Dye Glasses: A Guideline Toward the Fabrication of Amorphous Dye Films with Monomeric Absorption and Emission

Clara Schäfer, Sandra Hultmark, Yizhou Yang, Christian Müller, and Karl Börjesson\*



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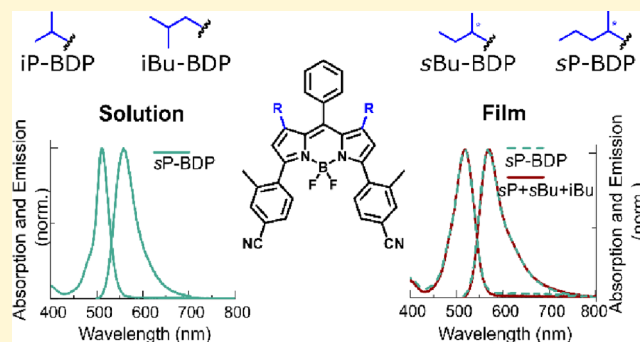
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**ABSTRACT:** The morphology of films containing photoactive materials is crucial for the performance of solid-state dye applications. Organic dyes tend to crystallize due to their usually planar molecular structure and the resulting intermolecular interactions. This leads to inhomogeneous films with crystalline, aggregated, and amorphous regions, decreasing device efficiency and complicating spectral analysis. Improving the glass-forming ability of organic dyes therefore presents a major challenge for solid-state dye applications. Here, we present a guideline to create organic dye glasses using BODIPY as a model dye. The method is based on the strategic design of BODIPY derivatives, equipped with short alkyl chains, in combination with blending of two or more derivatives. Mixing increases the entropy of the liquid state and lowers the thermodynamic driving force for crystallization as well as the kinetic fragility of the system. This enables the fabrication of homogeneous thin films without any additives. In these films, the dye molecules are trapped in a glassy state, featuring monomeric absorption and emission. This strategy leads to a BODIPY material with an amorphous character in thin films, dropcast films, and bulk. Further, the strategy is based on thermodynamics and is therefore expected to be general, enabling the transformation of any dye molecule into a glass former.



## INTRODUCTION

The conversion of a material into an amorphous solid, a glass, is called vitrification. The ability of materials to vitrify is of tremendous importance for amorphous pharmaceutical formulations,<sup>1,2</sup> vapor-deposited glasses used in organic light-emitting diodes (OLEDs),<sup>3–5</sup> and also stable organic solar cells.<sup>6–9</sup> There is therefore a consistent high demand for designing and engineering of materials having a good glass forming ability (GFA). For photoactive materials, controlling the nanostructure of thin films is crucial for solid-state applications.<sup>10,11</sup> The nanostructure of a thin film highly affects the photophysical properties and with that the device performance.<sup>12</sup> Crystallization and aggregation lead to unwanted side effects like scattering or undesired excimer emission, and therefore a great deal of research has been devoted to controlling the thin film nanostructure and improving the GFA of photoactive materials.<sup>13,14</sup> Organic dyes commonly comprise a flat, aromatic core. This structural feature causes them to crystallize, promoted by intermolecular  $\pi$ - $\pi$ -interactions. Strategies to reduce these interactions and prevent organic dyes from crystallizing include alkyl chain engineering, the use of additives like polymers that are mixed into the material, or by copolymerizing the dye to receive an

amorphous material.<sup>6,8,15–18</sup> However, these methods by themselves have drawbacks, causing limitations in applicability.

To reduce  $\pi$ - $\pi$ -interactions between molecules, alkylation of the  $\pi$ -conjugated system is a frequently used strategy. Here, a balanced interplay between the van der Waals interactions of the alkyl chains and the  $\pi$ - $\pi$ -interactions of the core is targeted.<sup>17</sup> This strategy is mainly used in the fields of soft materials,<sup>19</sup> functional molecular liquids,<sup>20</sup> or liquid crystals.<sup>21</sup> It is used to control crystallization,<sup>16,22</sup> and can also lead to unperturbed photophysical features in films by enclosing the  $\pi$ -core to shield it from its environment.<sup>20</sup> The attached alkyl chains although often surpass the  $\pi$ -core in molecular weight, making this strategy only applicable for concentration-independent applications.

To what extent a material is inclined to vitrify instead of crystallizing depends on the thermodynamic and kinetic factors that influence the rate of crystal nucleation as well as crystal

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growth.<sup>9,23,24</sup> To promote the glass formation of materials, one method that can be applied is the use of multicomponent mixtures. Mixing increases the entropy of the liquid state and therefore leads to a reduction of the thermodynamic driving force for crystallisation.<sup>25–27</sup> These mixtures contain several different components, which together form a glass. One or several of these components are often sugars or polymers.<sup>2,9</sup> This leads to limitations regarding the concentration of the active component in a similar manner as alkyl chain engineering. When different optically active components are mixed, the blend will adopt the combined photophysical properties of all the components, making monomeric absorption inaccessible.<sup>28</sup> These limitations need to be overcome to successfully make a glass having monomeric absorption and emission features.

Here we present a guideline on how to fabricate a dye glass having monomer-like absorption and emission properties as a thin film. We use BODIPY dyes, which have received a considerable amount of attention in various fields due to their versatile photophysical properties. They possess large molar absorptivity coefficients, fluorescence quantum yields close to unity, sharp absorption and emission bands, and high photostability.<sup>29,30</sup> The synthesis of these dyes is relatively straightforward, which makes functionalization and modification, and with this tuning of the photophysical properties possible.<sup>31–35</sup> Due to these properties, BODIPY dyes and their derivatives are used in a wide array of fields such as biological imaging,<sup>36–39</sup> sensing,<sup>40–42</sup> potential agents for photodynamic and photothermal therapy,<sup>43–45</sup> solar cells,<sup>46–49</sup> and optoelectronics.<sup>50–53</sup>

In this work, we will show that a good glass forming ability at room temperature is achieved for BODIPYs when combining alkyl chain engineering and mixing several components. By adding short alkyl chains to the  $\pi$ -core of the molecule, the photophysical properties of the dye remain unchanged.<sup>54</sup> This allows for mixing of several derivatives, differing only in the alkyl substituents, without broadening the absorption and emission envelope of the processed material. Therefore, by first choosing appropriate BODIPY derivatives and then mixing them, a glassy film is obtained at room temperature. The presented strategy is based on thermodynamics and is therefore expected to be general, enabling the transformation of any dye molecule into a glass former.

## ■ EXPERIMENTAL SECTION

**Synthesis.** All reactions were carried out under ambient conditions unless stated differently, for example, performed under a N<sub>2</sub> atmosphere. Glassware was oven dried prior to use unless indicated otherwise. Common reagents, solvents, or materials were obtained from Sigma-Aldrich Chemical Co. and used without further purification. Dry solvents for reactions sensitive to moisture and/or oxygen were obtained through a solvent purifying system (inert PureSolv-MD-5). Column chromatography was performed using silica gel (VWR 40 to 63  $\mu$ m) unless stated otherwise. Flash chromatography was performed on a Teledyne CombiFlash EZ prep using RediSepRf columns, normal-phase silica with a mesh size of 230 to 400, a particle size of 40 to 63  $\mu$ m, and a pore size of 60 Å unless stated otherwise. <sup>1</sup>H (<sup>13</sup>C) NMR spectra were recorded on a Varian 400 spectrometer (400 MHz <sup>1</sup>H; 100 MHz <sup>13</sup>C) at room temperature using CDCl<sub>3</sub> (the final products containing tetramethylsilane with 0.00 ppm as an internal reference) or DMSO-*d*<sub>6</sub> as solvent. Coupling constants (*J* values) are given in Hertz (Hz) and chemical shifts are reported in parts per million (ppm). <sup>13</sup>C spectra are decoupled from <sup>1</sup>H. Low-resolution MS was obtained from a GC/MSD system from Agilent 7820A GC System in tandem with an

Agilent 5977E Mass Spectrometer. High-resolution MS was obtained from an Agilent 1290 infinity LC system equipped with an autosampler in tandem with an Agilent 6520 Accurate Mass Q-TOF LC/MS. Melting points were measured using a BÜCHI Melting Point B-545 instrument. IR spectra were recorded using an INVENIO R FT-IR from BRUKER.

The full synthesis is described in the [Supporting Information](#). The Boc-protected Suzuki-coupling products were synthesized following a previously reported procedure.<sup>55</sup> To obtain differently alkylated compounds, the starting materials were changed accordingly.

**General procedure for BODIPY Synthesis.** Boc-protected Suzuki-coupling product (2.00 equiv) was added to the reaction vessel, which was subsequently evacuated and refilled with N<sub>2</sub> (3 cycles). The starting material was dissolved in dry THF (5 mL·mmol<sup>-1</sup>). NaOMe (25 w % in MeOH, 0.4 mL·mmol<sup>-1</sup>) was added to the reaction mixture and the reaction was stirred until full conversion of the starting material (approximately 2 h). H<sub>2</sub>O was added to the reaction mixture and the phases were separated. The aqueous phase was extracted with diethyl ether (2 × 20 mL) and the combined organic phases were washed with brine (25 mL). The solvent was subsequently removed under reduced pressure. The product H-SCP-NH was obtained as a light-yellow oil and was directly used without further purification or analysis.

The deprotected Suzuki-coupling product (2.00 equiv) and benzaldehyde (1.00 equiv) were dissolved in DCM<sub>dry</sub> (100 mL·mmol<sup>-1</sup> benzaldehyde<sup>-1</sup>). A catalytic amount of TFA was added to the solution and the solution was stirred until TLC showed full conversion of benzaldehyde (TLC: SiO<sub>2</sub> hexane/EtOAc 10% ca. 14 h). After full conversion was determined, DDQ (2.05 equiv) in DCM (10 mL·mmol<sup>-1</sup> DDQ<sup>-1</sup>) was added to the reaction mixture. The reaction mixture, which changed color from dark pink to dark purple, was stirred at room temperature for 1 h. Et<sub>3</sub>N (6.5 mL·mmol<sup>-1</sup> benzaldehyde<sup>-1</sup>) was added to the reaction mixture, accompanied by a color change to dark red, and stirred for 30 min at room temperature. Afterward BF<sub>3</sub>·OEt<sub>2</sub> (6.5 mL·mmol<sup>-1</sup> benzaldehyde<sup>-1</sup>) was added to the reaction mixture and the now dark purple solution was stirred overnight. H<sub>2</sub>O was added to the reaction mixture, and the two-phase mixture was stirred for 1 h. Thereafter the phases were separated, and the organic phase was washed with H<sub>2</sub>O (3 × 200 mL), afterward dried over Na<sub>2</sub>SO<sub>4</sub>, and then the solvent was removed under reduced pressure. The crude product was purified by column chromatography as specified for the single products.

**sBu-BDP.** The product was prepared following the general procedure for BODIPY synthesis and purified using column chromatography (1st: SiO<sub>2</sub> column, eluent: hexane/EtOAc 0–50%; 2nd: SiO<sub>2</sub> column, eluent: hexane/EtOAc 20%). The product was obtained as a red crystalline solid (1.52 g, 2.49 mmol, 83%), starting from tert-butyl 4-(*sec*-butyl)-2-(4-cyano-2-methylphenyl)-1H-pyrrole-1-carboxylate (6.00 mmol, 2.03 g, 2 equiv.).

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>):  $\delta$  7.60–7.47 (m, 5H), 7.50–7.40 (m, 6H), 6.17 (s, 2H), 2.24 (s, 6H), 1.39–1.26 (m, 4H), 1.23–1.12 (m, 2H), 0.84 (dd, *J* = 6.5, 3.7 Hz, 6H), 0.56 (td, *J* = 7.1, 4.4 Hz, 6H). <sup>13</sup>C NMR (101 MHz, CDCl<sub>3</sub>):  $\delta$  156.3, 154.9, 145.7, 138.3, 137.6, 134.3, 133.1, 132.0, 130.7, 129.6, 128.8, 128.6, 128.4, 128.3, 118.8, 112.5, 32.8, 32.7, 31.4, 31.3, 22.2, 22.1, 20.1, 12.2. IR:  $\nu_{\max}$ /cm<sup>-1</sup> 2962, 2926, 2871, 2230, 1521, 1476, 1405, 1314, 1167, 1147, 1125, 1102, 1039, 829, 753. HRMS: (ESI<sup>+</sup>) *m/z*: calcd. For (M + H)<sup>+</sup> C<sub>39</sub>H<sub>38</sub>BF<sub>2</sub>N<sub>4</sub>: 611.31576; found: 611.3151. mp 266 °C.

**iBu-BDP.** The product was prepared following the general procedure for BODIPY synthesis and purified using column chromatography (1st: SiO<sub>2</sub> column, eluent: hexane/EtOAc 10–20%; 2nd: SiO<sub>2</sub> column, eluent: pentane/EtOAc 10%). The product was obtained as a red crystalline solid (1.27 g, 2.08 mmol, 69%), starting from tert-butyl 2-(4-cyano-2-methylphenyl)-4-isobutyl-1H-pyrrole-1-carboxylate (6.00 mmol, 2.03 g, 2 equiv.).

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>):  $\delta$  7.64–7.51 (m, 3H), 7.51–7.40 (m, 8H), 6.12 (s, 2H), 2.25 (s, 6H), 1.54 (d, *J* = 6.5 Hz, 3H), 1.45 (dq, *J* = 14.2, 6.7 Hz, 2H), 0.60 (d, *J* = 6.4 Hz, 12H). <sup>13</sup>C NMR (101 MHz, CDCl<sub>3</sub>):  $\delta$  154.4, 148.4, 145.7, 138.3, 137.5, 134.0, 133.1, 132.5, 130.7, 129.4, 128.8, 128.7, 128.6, 122.0, 118.8, 112.6, 37.0,

29.3, 22.3, 20.1. IR:  $\nu_{\max}/\text{cm}^{-1}$  2957, 2927, 2869, 2230, 1521, 1476, 1404, 1302, 1153, 1121, 1069, 833, 757, 722. HRMS: (ESI<sup>+</sup>)  $m/z$ : calcd. For (M + H)<sup>+</sup> C<sub>39</sub>H<sub>38</sub>BF<sub>2</sub>N<sub>4</sub>: 611.31576; found: 611.3150. mp 279 °C.

**sP-BDP.** The product was prepared following the general procedure for BODIPIY synthesis and purified using column chromatography (1st: SiO<sub>2</sub> column, eluent: hexane/EtOAc 10–20%; 2nd: SiO<sub>2</sub> column, eluent: hexane/EtOAc 5–15%). The product was isolated as a red crystalline solid (872 mg, 1.37 mmol, 46%, isolated yield), starting from tert-butyl 2-(4-cyano-2-methylphenyl)-4-(pentan-2-yl)-1H-pyrrole-1-carboxylate (6.00 mmol, 2.11 g, 2 equiv).

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>):  $\delta$  7.57–7.49 (m, 5H), 7.47–7.42 (m, 6H), 6.18 (s, 2H), 2.24 (s, 6H), 1.45 (h,  $J$  = 6.7 Hz, 2H), 1.33–1.21 (m, 2H), 1.20–1.09 (m, 1H), 1.08–0.97 (m, 1H), 0.83 (t,  $J$  = 6.7 Hz, 6H), 0.70 (td,  $J$  = 7.3, 5.0 Hz, 6H). <sup>13</sup>C NMR (101 MHz, CDCl<sub>3</sub>):  $\delta$  156.5, 154.9, 145.6, 138.3, 137.6, 134.3, 133.1, 130.7, 129.6, 129.6, 128.6, 128.5, 128.4, 128.3, 118.8, 112.5, 40.7, 40.5, 31.1, 22.7, 22.3, 20.8, 20.1, 14.1. IR:  $\nu_{\max}/\text{cm}^{-1}$  2958, 2927, 2869, 2230, 1521, 1475, 1405, 1309, 1147, 1125, 1098, 829, 757. HRMS: (ESI<sup>+</sup>)  $m/z$ : calcd. For (M + H)<sup>+</sup> C<sub>41</sub>H<sub>42</sub>BF<sub>2</sub>N<sub>4</sub>: 639.34706; found: 639.3469. mp 132 °C.

**iP-BDP.** The product was prepared following the general procedure for BODIPIY synthesis and purified using column chromatography (1<sup>st</sup>: SiO<sub>2</sub> column, eluent: hexane/EtOAc 0–50%; 2<sup>nd</sup>: SiO<sub>2</sub> column, eluent: DCM). The product was obtained as a red crystalline solid (1.27 g, 2.18 mmol, 73%), starting from tert-butyl 2-(4-cyano-2-methylphenyl)-4-isopropyl-1H-pyrrole-1-carboxylate (6.00 mmol, 1.95 g, 2 equiv).

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>):  $\delta$  7.58–7.48 (m, 5H), 7.48–7.41 (m, 6H), 6.22 (s, 2H), 2.24 (s, 6H), 1.62 (p,  $J$  = 6.7 Hz, 2H), 0.87 (d,  $J$  = 6.7 Hz, 12H). <sup>13</sup>C NMR (101 MHz, CDCl<sub>3</sub>):  $\delta$  157.2, 154.9, 145.7, 138.3, 137.6, 134.4, 133.1, 131.4, 130.7, 129.7, 128.6, 128.4, 118.8, 118.5, 112.5, 25.9, 24.5, 20.1. IR:  $\nu_{\max}/\text{cm}^{-1}$  2969, 2928, 2871, 2229, 1519, 1477, 1405, 1314, 1169, 1147, 1126, 1089, 1065, 1032, 963, 829, 754. HRMS: (ESI<sup>+</sup>)  $m/z$ : calcd. For (M + H)<sup>+</sup> C<sub>37</sub>H<sub>34</sub>BF<sub>2</sub>N<sub>4</sub>: 583.28446; found: 583.2845. mp 339 °C.

**Sample Preparation.** Glass substrates for films (25 × 25 mm) were precleaned by sonication for 1 h in an alkaline solution (0.5% of Hellmanex in distilled water), then rinsed with water and sonicated for 1 h in water and ethanol, respectively. To avoid any inner filter effects, which are dependent on the thickness of the films, the neat films as well as the blends were prepared to be very thin (~10 nm). For thin neat films, solutions of the BODIPIY dyes ( $c_{\text{dye}} = 0.5 \text{ mg mL}^{-1}$ ) in DCM were spin-coated (45 s, R.T., 1500 rpm) (Laurell) on the glass substrates. For thin blended films, solutions of the BODIPIY dyes ( $c_{\text{dye}} = 0.5 \text{ mg mL}^{-1}$ ) in DCM were spin-coated (45 s, R.T., 1500 rpm) (Laurell) on the glass substrates. In the blends, the ratio of the components used is 1:1 or 1:1:1. The dyes were weighed in as powders, dissolved, and the solution was stirred for 15 min prior to spin coating. For dropcasted films, solutions of the BODIPIY dyes ( $c_{\text{dye}} = 10.0 \text{ mg mL}^{-1}$ ) in DCM (0.5 mL) were dropcasted onto a precleaned glass substrate.

**Scanning Calorimetry.** Differential scanning calorimetry (DSC). Measurements were performed using a Mettler Toledo DSC2 equipped with a Gas controller GC 200 with a heating and cooling rate of 0.17 K s<sup>-1</sup> (10 °C min<sup>-1</sup>). Around 2–5 mg of materials was collected into 40  $\mu\text{L}$  DSC Al crucibles. Fast scanning calorimetry (FSC). Measurements were carried out using a Mettler Toledo Flash DSC 1. Solutions were prepared and drop cast on glass slides. A small amount of material was transferred from the glass slide to the FSC sensor. Samples were first heated to above  $T_m$  for each material and then cooled to 25 °C with cooling rates ranging from -0.1 to -1000 K s<sup>-1</sup>, followed by heating again at 2000 K s<sup>-1</sup>. Two different methods were used to calculate the fictive temperatures,  $T_f'$ , for the fragility plot (Figure S4). Moynihan's method was used if  $T_f'$  was above the  $T_{g,\text{onset}}$  and it is calculated according to

$$\int_{T_f'}^{T_g} (C_{\text{pl}} - C_{\text{pg}}) dT = \int_{T_g}^{T_g} (C_{\text{p}} - C_{\text{pg}}) dT$$

where  $C_{\text{pl}}$  and  $C_{\text{pg}}$  are the heat capacities of the liquid and the glass, respectively, and  $C_{\text{p}}$  is the apparent heat capacity of the sample. If  $T_f'$  instead was below the  $T_{g,\text{onset}}$ , a simplified extrapolation version was used:  $\int_{T_f'}^{T_g} (C_{\text{pl}} - C_{\text{pg}}) dT = 0$ .

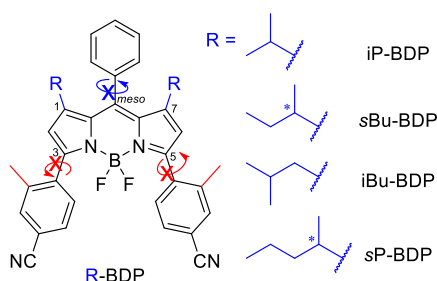
**Optical Spectroscopy.** Absorption spectra were measured using a spectrophotometer (LAMBDA 950, PerkinElmer). Steady-state emission spectra, excitation spectra, and emission lifetimes were measured with a spectrofluorometer (FLS1000, Edinburgh Instrument) and are corrected using the emission correction files provided by the manufacturer. The emission spectra were first converted to the energy scale using the lambda-squared correction for the emission maximum when assessing the Stokes shift. The fluorescence quantum yields of samples in solution were calculated using the relative method, using a standard (fluorescein) with a known emission quantum yield.<sup>56</sup> Fluorescein in 0.1 M NaOH ( $\Phi_f = 0.91$ ) was used as a reference compound for  $\Phi_f$  determination. (Excitation at 480 nm, refractive index: 1.33). The samples were measured in toluene solution at 22 °C, refractive index: 1.497. For the emission lifetime measurements, the samples were excited by a 510 nm picosecond pulsed diode laser (Edinburgh Instruments). The internal response function (IRF) was measured using a soap and water solution having a high scattering effect for samples in solution, and an Al mirror for films. Emission quantum yields of the solid samples were measured using the spectrofluorometer (FLS1000, Edinburgh Instrument) equipped with an integrating sphere. As a reference, a blank glass substrate was used.

**Optical Microscopy.** Optical micrographs were recorded with a Zeiss AxioScope 5 equipped with a pair of crossed polarizers.

**Grazing-Incidence Wide-angle X-ray Scattering (GIWAXS).** GIWAXS data were obtained on a Mat/Nordic SAXSLAB instrument. The detector used was a Pilatus3 300K R from Dectris, and the source was a Rigaku Micromax-003 with a Cu target. All measurements were carried out at a pressure of approximately 0.1 mbar in the entire flightpath. The sample to detector distance was set to 127 mm, and the incidence angle was set to 0.2°. Exposure times were 5 min. Before each measurement, a calibration of the sample tilt was performed with a z-scan and a rocking curve.

## RESULTS AND DISCUSSION

**Molecular Design and Synthesis.** To fabricate a dye glass, the first thing one should pay careful consideration to is the design of the molecules used in the film. As mentioned previously, to prevent crystallization,  $\pi$ - $\pi$ -interactions must be minimized or eliminated. Adding alkyl chains to the core structure is a commonly used strategy to counteract these  $\pi$ - $\pi$ -interactions.<sup>17,21</sup> Branched alkyl chains have been shown to be more efficient in increasing solubility and interfering with aggregation.<sup>54,57,58</sup> In our molecular design, we decided to attach small but relatively bulky side chains at the 1,7-positions of the BODIPIY core, shown in Figure 1. We chose branched alkyl chains of increasing length, isopropyl (iP), isobutyl (iBu), *sec*-butyl (sBu) and *sec*-pentyl (sP). The sBu and sP chains, in addition to being relatively bulky, introduce a stereogenic center, which means a racemic mixture of the BODIPIY will be obtained. A mixture of isomers will already increase the entropy in the system compared to enantiopure compounds and will therefore be beneficial to prevent aggregation.<sup>54</sup> Paracyanophenyl groups at positions 3,5 were attached in order for the compounds to be appropriate for thermal analysis. This is important since the molecule sublimates before melting if the molecular weight is too low, which we observed for compounds having only the phenyl rings without a -CN group (Figure S1). Furthermore, ortho-methyl groups on the



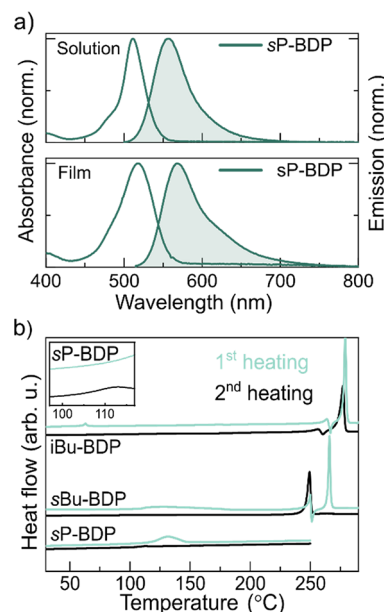
**Figure 1.** Structure of the designed and synthesized BODIPY derivatives, with the different alkyl chains shown in blue.

para-cyanophenyl rings in the 3,5-positions prevent planarization by blocking the rotation around the C–C bond, as indicated in Figure 1, and therefore make the molecule less prone to crystallize. The phenyl group on the meso-methine bridge was added for synthetic convenience. The planarization of this ring is blocked by the small alkyl chains in the 1,7-positions, which further reduces the crystallization potential of the molecules.

The different BODIPY derivatives were synthesized from pyrroles, carrying different alkyl chains at the 3-position, and benzaldehyde. The alkylated pyrroles were obtained via a Wittig reaction of triethyl phosphonoacetate and the corresponding aldehyde. This was followed by a Van-Leusen reaction using *p*-toluenesulfonylmethyl isocyanide (TosMIC) to form the alkylated pyrrole carboxylate esters and subsequently decarboxylation to the alkylated pyrroles. The boronic acids of the alkylated pyrroles were synthesized and coupled to 4-bromo-3-methylbenzotrile via a Suzuki-Miyaura coupling reaction.<sup>55</sup> To obtain the BODIPYs, previously reported synthetic strategies were followed.<sup>59,60</sup> The synthesized pyrroles, together with benzaldehyde, were first submerged to condensation in the presence of a catalytic amount of trifluoroacetic acid to form the dipyrromethane. This was subsequently treated with the oxidant 2,3-dichloro-5,6-dicyano-*p*-benzoquinone (DDQ) to form the dipyrromethene, and finally, Et<sub>3</sub>N followed by BF<sub>3</sub>·OEt<sub>2</sub> were added for complexation to finalize the BODIPY structure.

**Neat Materials in Solution and Solid State.** Following the synthesis, the derivatives were first analyzed spectroscopically to obtain information about their photophysical properties. The BODIPY derivatives can be expected to have similar photophysical properties since the only structural changes

within them are the alkyl chains attached to the core structure. The derivatives in solution show as predicted highly similar absorption and emission features as shown in Table 1, Figures 2a and S7, with absorption and emission maxima at 511 and



**Figure 2.** (a) BODIPY absorbance and emission in solution and solid state, here represented by sP-BDP (see Figures S7 and S8a,b for other BODIPY derivatives) (b) differential scanning calorimetry (DSC) first (green) and second (black) heating thermograms of iBu-BDP, sBu-BDP and sP-BDP measured at a heating rate of 0.17 K s<sup>-1</sup>. With a zoom on the T<sub>g</sub> of sP-BDP at 110 °C in the second heating curve.

558 nm, respectively, and molar absorptivity coefficients of 61.000 to 72.000 M<sup>-1</sup> cm<sup>-1</sup>. The fluorescence quantum yields (Φ<sub>f</sub>) are lower than those common for BODIPYs. The sBu-BDP without the phenyl group in the meso-position is reported to have a close to unity emission quantum yield.<sup>55</sup> A reduced emission quantum yield through a rotating phenyl substituent in the meso position has been described previously.<sup>60,61</sup> Yet, methyl groups at the 1,7-positions are commonly sufficient to block the free rotation. Therefore, the drop in Φ<sub>f</sub> is unexpected and can only be attributed to the movement of the alkyl chains and phenyl group that seemingly affect the non-radiative pathway in different amounts. For

**Table 1. Spectroscopic Data of the BODIPY Derivatives in Toluene Solution at Room Temperature, and of Thin Films of the BODIPY Derivatives<sup>a</sup>**

	sBu-BDP		sP-BDP		iBu-BDP		iP-BDP	
	solution	Film	solution	film	solution	film	solution	film
λ <sub>max</sub> Abs [nm]	511	518	511	518	513	519	511	518
λ <sub>max</sub> Em [nm]	558	568	558	568	559	571	558	572
λ <sub>max</sub> Exc [nm]	512	515	512	517	513	517	512	516
Stokes shift [cm <sup>-1</sup> ]	1646	1699	1646	1755	1604	1755	1646	1822
ε [M <sup>-1</sup> cm <sup>-1</sup> ]	66600		68500		61500		72000	
Φ <sub>f</sub>	0.29 <sup>b</sup>	0.10 <sup>c</sup>	0.33 <sup>b</sup>	0.11 <sup>c</sup>	0.48 <sup>b</sup>	0.07 <sup>c</sup>	0.15 <sup>b</sup>	0.07 <sup>c</sup>
τ <sub>f</sub> [ns]	1.45	2.04 (0.995) <sup>d</sup>	1.52	1.98 (0.993) <sup>d</sup>	2.27	2.26 (0.991) <sup>d</sup>	0.76	0.86 (0.962) <sup>d</sup>
τ <sub>f-2</sub> [ns]		10.38 (0.005) <sup>d</sup>		9.19 (0.007) <sup>d</sup>		12.11 (0.009) <sup>d</sup>		5.31 (0.038) <sup>d</sup>

<sup>a</sup>The films were spin coated from DCM (0.5 g L<sup>-1</sup>). <sup>b</sup>Fluorescein in 0.1 M NaOH (Φ<sub>f</sub> = 0.91) was used as a reference compound for Φ<sub>f</sub> determination.<sup>56</sup> (Excitation at 491 nm, Refractive index: 1.33). <sup>c</sup>Quantum yields of the films were measured using an integrating sphere. <sup>d</sup>Pre-exponential factors.

future solid-state studies, the divergent fluorescence lifetimes are, however, not of importance since movements such as rotations are likely blocked in the solid state. Thus, to summarize, the absorption and emission features of the four derivatives are close to indistinguishable. This makes them suitable candidates for mixing, which will be discussed later.

To explore how the single derivatives behave in the solid state, thin films of the dyes were fabricated (thickness  $\sim 10$  nm) and the photophysical properties were evaluated. These thin films were fabricated by spincoating of a dilute solution of the dye in DCM ( $0.5 \text{ g L}^{-1}$ ). The spectroscopic evaluation of the thin films is summarized in Table 1 and Figures S8a,b, S12, and S14. The thin films exhibit monomeric absorption and emission envelopes, as shown in Figure 2a, with a slight shift in absorption and emission maxima of 7–8 and 10–14 nm, respectively, in comparison to the dyes in toluene solution. This shift can be attributed to the change in environment of the dye molecules in the solid state as compared to the solution state. A polarity dependence of the absorbance maximum can also be observed when the sP-BDP is measured in different solvents (Figure S9). The fluorescence quantum yields drop to 7–11%. Instead of one distinct fluorescence lifetime, there is now a second, longer lifetime detected, although this one only contributes to the emission to a minor extent. The films are homogeneous and show no sign of crystallization or aggregation when analyzed with an optical microscope equipped with cross polarizers. This glass state of the films could be a result of their being too thin to provide a sufficient area density of nucleation sites for the material to crystallize. Alternatively, the glass state can be a result of the processing method. Spincoating leads to the rapid removal of the processing solvent, preventing the crystallization of the material. Anyhow, the conditions for crystallization are absent, and the molecules can therefore be considered kinetically trapped in a glass state. This would also explain the broadening of the absorption. It is most likely due to the conformational restrictions of the dyes in the film. Observations of the iP-BDP thin film over several days show an abrupt change in both absorption and emission features (Figure S10a). The very distinct broadening of the absorption peak as well as the emerging of a second emission band at around 620 nm indicate that aggregation of the molecules in the thin film occurs over time in a cooperative manner.

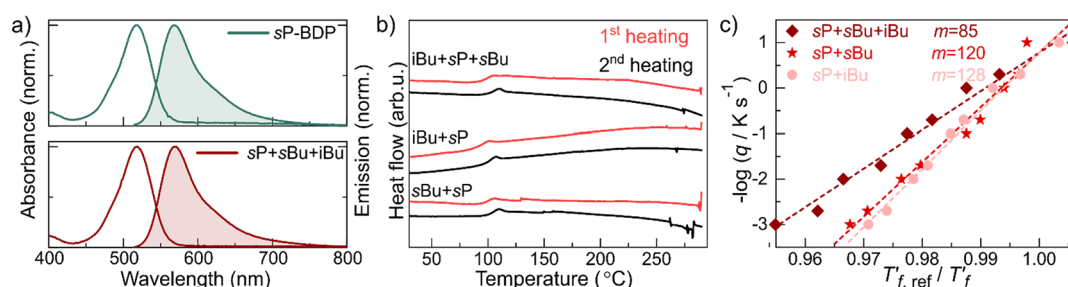
To further evaluate the crystallization behavior of the single materials, thick films (about 3–4  $\mu\text{m}$ ) were made by dropcasting. In this case, the films have both more material, which should provide a higher area density of nucleation sites, and dry slower, resulting in a longer period of time during which crystallites can form. This leads to the films being less homogeneous and less smooth, yet provides the opportunity of investigating the GFA of the materials in a scenario where the molecules are less prone to get kinetically trapped. The thick films of the BODIPY derivatives show birefringent areas using polarized optical microscopy for the iP-BDP and iBu-BDP films, which indicates the presence of ordered domains (Figure S16a,b). The iP-BDP film showed clear peaks in Grazing-Incidence Wide-Angle X-ray Scattering (GIWAXS) analysis, indicating multicrystallinity of the film (Figure S18a). The sBu-BDP and sP-BDP films, however, are more homogeneous and show no distinct birefringent areas (Figure S16c,d). This can be explained by either the ordered domains being very small, or by the nature of the alkyl chains. As mentioned before, both the sBu and the sP chains introduce a stereogenic center into

the molecule, leading to the formation of a racemic mixture. A mixture of isomers increases the entropy of the system compared to the enantiopure system. Thus, it becomes more difficult to crystallize. Amorphous dropcasted films have already been observed for a racemic mixture of sterically hindered BODIPY derivatives carrying paracyclophane units in the 3,5-positions.<sup>62</sup> The film formed by these derivatives was described as “not homogeneous with spots or dried droplets”. Here in comparison, our films do not dewet from the surface, and in the case of sBu-BDP and sP-BDP, transparent films without birefringent areas are formed, as can be seen in optical micrographs.

The thick films of the racemic BODIPY mixtures did not show birefringent areas, which could indicate that the molecules do not aggregate or crystallize. Typically, the crystal nucleation rate peaks close to the glass transition temperature ( $T_g$ ), while the growth rate peaks close to the melting temperature ( $T_m$ ). As a consequence, both rates are highly affected by the temperature dependence of the viscosity above  $T_g$  and therefore the cooling rate. How the viscosity of a liquid increases upon cooling at the  $T_g$  is described by its kinetic fragility ( $m$ ).<sup>63,64</sup> This measure is also used to classify glass-formers into fragile and strong glass-formers.<sup>65</sup>

To see if the neat materials are amorphous, or can be amorphous under other processing conditions, the thermal behavior of the material needs to be measured in bulk. This will also give additional information about the glass forming abilities of the material. First, to get information about the thermal behavior of the compounds during heating, differential scanning calorimetry (DSC) was employed. IP-BDP was determined not to be appropriate for this study. Thermogravimetric analysis showed a weight loss of material in direct proximity to the melting temperature (Figure S2), and therefore, a loss of mother material cannot be avoided during the melting of this compound. Both, iBu-BDP and sBu-BDP show a high  $T_m$  with peaks at 279 and 266  $^\circ\text{C}$ , respectively (Figure 2b). A melting peak can be seen in both heating cycles, which indicates that the materials are crystallizing again after melting. Compared to these two derivatives, sP-BDP shows a much lower  $T_m$  with a maximum at 132  $^\circ\text{C}$ . In addition to the lower  $T_m$ , melting can only be seen in the 1st but not in the 2nd heating thermogram. Instead, a glass transition in the second heating thermogram was observed with a  $T_g$  around 110  $^\circ\text{C}$ . The disappearance of a  $T_m$  in combination with the appearance of a  $T_g$  indicates the absence of crystallization when cooling down, thus the formation of an amorphous solid state.

To further investigate the GFA and determine its fragility, fast scanning calorimetry (FSC) was employed. Several different processes were observed within the same heating and cooling cycle, including glass transition and crystallization (Figure S5), preventing a detailed analysis of the fragility. The single derivatives are fragile and crystallize even if cooled quickly at a cooling rate of  $-100 \text{ K s}^{-1}$ . To summarize, when the single material films are very thin ( $\sim 10$  nm), the molecules are trapped in a glass state. Instead, for films with a thickness of about 3–4  $\mu\text{m}$ , the tendency to crystallize increases. This was seen for iP and iBu-BDP, which show birefringent areas in the thick but not in thin films. Having a racemic mixture like in the case of sBu-BDP and sP-BDP, the entropy of the system is increased, leading to less crystalline thick films. DSC measurements of the materials in bulk show a decrease in  $T_m$  for iBu-BDP, sBu-BDP, and sP-BDP compared to iP-BDP. For sP-BDP, the absence of a  $T_m$  in combination with a  $T_g$  at



**Figure 3.** (a) Absorbance and emission for a sP-BDP film (top), representative for all neat BODIPIY derivatives and the film of the ternary blend sP + sBu + iBu (bottom), representative for all blends (Absorptivity at 518 nm =  $170 \pm 20 \times 10^3 \text{ cm}^{-1}$ ). (b) DSC first (red) and second (black) heating thermograms of iBu + sP + sBu, iBu + sP and sBu + sP measured at  $q = 0.17 \text{ K s}^{-1}$  (c) Fragility plot with  $-\log q$  versus  $T'_{f, \text{ref}}/T'_f$  of sP + sBu + iBu (dark red squares), sP + sBu (red stars), and sP + iBu (light red circles).  $T'_f$  is determined from fast scanning calorimetry (FSC) with the cooling rate  $q$  ranging from  $-0.1$  to  $-1000 \text{ K s}^{-1}$  and  $T'_{f, \text{ref}}$  is measured with DSC at  $q = 0.17 \text{ K s}^{-1}$ .

110 °C in the second heating thermogram indicates the formation of a glassy material.

**Mixing.** In applications that require an amorphous material, the glass forming ability should be as high as possible. How mixing impacts the thermodynamic driving force for crystallization depends on the change of free energy  $\Delta G_{\text{mix}} = \Delta H_{\text{mix}} - T\Delta S_{\text{mix}}$  and therefore, several factors. First, the change in enthalpy upon mixing ( $\Delta H_{\text{mix}}$ ) has to be considered. Attractive molecular interactions between the different components, like  $\pi$ - $\pi$ -interactions, can result in  $\Delta H_{\text{mix}} < 0$ , which can lead to a lower free energy of the mixed state and will result in a higher crystallization rate. Whereas when  $\Delta H_{\text{mix}} > 0$ , the free energy of the mixed state will increase and at some point, if  $\Delta H_{\text{mix}} > T\Delta S_{\text{mix}}$ , will not favor the mixing of the different components.<sup>26</sup> The type of components used is, therefore, an important factor when considering blending. Another factor is the number of different components. This naturally affects the entropy of the system upon mixing ( $\Delta S_{\text{mix}}$ ) and has also been shown to affect the kinetic fragility of multicomponent mixtures.<sup>64</sup> In the system used in this work, the components are designed to have highly similar molecular structures. Hence, this system can be considered close to ideal, and  $\Delta H_{\text{mix}}$  can therefore be assumed to be close to 0. Consequently, the change of free energy  $\Delta G_{\text{mix}} = \Delta H_{\text{mix}} - T\Delta S_{\text{mix}}$  depends on the change in entropy ( $\Delta S_{\text{mix}}$ ) of the system upon mixing.

Blending two or more of the BODIPIY derivatives further increases the entropy, as compared to the racemic mixtures of sP- and sBu-BDP, which should decrease the ability of the system to crystallize. To examine if this is applicable to our system, three binary blends and one ternary blend were made. The binary blends sBu + iBu, sP + sBu, and sP + iBu as well as the ternary blend sP + sBu + iBu were made and analyzed in the same fashion as the films of neat derivatives. The DSC heating thermogram of the sBu + iBu blend showed several different processes, including a glass transition at 110 °C, crystallization and melting and was therefore excluded from the study (Figure S3). All the other three blends show only one clear transition in the DSC measurements, which will be analyzed in detail further down in this section. As for the neat materials, thin films of the blends were processed by spincoating. The absorption as well as the emission transitions are of monomeric nature and are indistinguishable from the single component films as shown in Figures 3a (Figure S8c,d and Table S1). Observations of the ternary blend over time show no change in the absorption envelope. Furthermore, the emission shows the development of a small shoulder at around

620 nm over the course of several days (Figure S10b,d). As for the single component thin films, no aggregation can be observed under a polarized optical microscope. Furthermore, and importantly, the thick films (3–4  $\mu\text{m}$ ) fabricated by dropcasting also show no sign of birefringence (Figure S17). GIWAXS analysis shows no sign of crystallinity for the ternary blend, whereas the thick film of iP-BDP shows clear peaks (Figure S18b). This leads to the conclusion that by the mixing of BODIPIY derivatives an amorphous material can be produced.

To determine if the blending of derivatives improved the GFA, first DSC was measured in bulk. The 1st and 2nd DSC heating thermograms, shown in Figure 3b, feature clear glass transitions. The  $T_g$  for all the four prepared blends has a similar value of 110 °C, and can therefore not be attributed to any specific derivative. No obvious change in the  $T_g$  can be observed by blending two or three derivatives. More noteworthy is that no melting was observed in the first heating thermograms of either of the blends. This leads to the conclusion that the material is unable to crystallize or aggregate as a result of blending the derivatives at room temperature.

Thus, it is not necessary to heat the material above the melting point in order to form an amorphous solid state, like seen in the DSC measurement of sP-BDP.

To determine the kinetic fragility of the blends, FSC was carried out (see Figure S4 for information on fragility calculations and Figure S6 for FSC data).<sup>64</sup> The binary blends have a relatively high fragility of  $m = 120$  and  $m = 128$  for sP + sBu and sP + iBu, respectively (Figure 3c). For the ternary blend, the fragility decreased to  $m = 85$ . The ternary blend has therefore a slightly better glass forming ability than the binary blends. Furthermore, in all cases the slope is constant, indicating that no liquid–liquid transition occurs. We have previously experienced the occurrence of liquid–liquid transitions in dye glasses, leading to low fragility but also aggregation and an associated change in the emission envelope.<sup>64</sup> In the present study, we see no tendency of a liquid–liquid transition in the FSC data nor an aggregation associated change of the emission. The optical properties of the single molecules are conserved after mixing (Figures 3a, S8c,d, S13, and S15).

In summary, making binary and ternary blends of the BODIPIY derivatives enable processing of homogeneous thin and thick films. The thin films of the multicomponent blends exhibit monomeric absorption and emission features, which are indistinguishable from the single-component thin films. DSC measurements of the blends show no melting features, which

indicates the creation of an amorphous solid even without erasing the thermal history by first heating over  $T_m$ . We also observed a decrease in the kinetic fragility from  $m = 128$  to  $m = 85$  by going from a binary to a ternary blend.

## CONCLUSIONS

We presented a guideline for how to make dye materials that form an amorphous solid state as well as homogeneous thin films that exhibit monomeric absorption and emission. First and foremost, the molecular design of the single components is crucial. To block free rotation of the 1,7-phenyl substituents and decrease the potential of intermolecular  $\pi$ - $\pi$  interactions, methyl groups were attached in the ortho-position. The free rotation of the meso-phenyl substituent was blocked by attaching alkyl chains in the 3,5-positions of the BODIPY core to prevent planarization and with that  $\pi$ - $\pi$  interactions. By attaching different alkyl chains at the 3,5-positions, derivatives of the BODIPY were synthesized, which exhibit almost indistinguishable absorption and emission properties. Adding branched alkyl chains that create a stereocenter, leading to racemic mixtures, decreases the crystallization potential of the single components.

Having discussed the molecular design, the next step is the processing of the films. Spincoating, due to the rapid removal of the solvent, leads to the formation of homogeneous glassy films where the molecules get kinetically trapped into an amorphous state and no crystallization occurs. Dropcasting offers more material as well as a slower removal of the solvent. The thick films, therefore, show aggregation for iP-BDP and iBu-BDP. For sBu-BDP, sP-BDP, both racemic mixtures, no birefringent areas were observed.

Finally, blending photophysically indistinguishable derivatives lowers the aggregation and crystallization tendency of the system. The blends prepared by spincoating result in homogeneous thin films having monomeric absorption and emission. Dropcasting of the same material results in homogeneous films without birefringent regions. The ternary blend also showed no sign of crystallinity using GIWAXS analysis. This, as well as DSC measurements, demonstrates the formation of an amorphous material without the necessity of the previous melting. In contrast, the neat materials all showed a clear  $T_m$ , and only sP-BDP showed a  $T_g$ . The fragility of the blends could be obtained from FSC. A decrease in the fragility of the ternary blend as compared to the binary ones was observed and the absence of a liquid-liquid transition was confirmed.

Following these steps, we prepared BODIPY materials that do not need additives to prevent crystallization and therefore solidify into a glass without any additives. This approach combines alkyl chain engineering, a method abundantly used to improve device performance and mixing of several components, to lower the crystallization tendency of the material. The here described strategy enables the formation of materials that have the potential to bring fields like optoelectronics, where dyes are used in the solid state and crystallization has so far been a limiting factor, forward.

## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.chemmater.2c02761>.

Synthesis of the BODIPY derivatives; Wittig reaction; Van-Leusen reaction; decarboxylation toward the alkylated pyrroles and Boc-protection; synthesis of pyrrole boronic acid; Suzuki-Miyaura coupling; thermogravimetric analysis; differential scanning calorimetry; fast scanning calorimetry; normalized absorbance, excitation, and emission of the BODIPY derivatives; fluorescence lifetimes; fluorescence quantum yields; optical micrographs; grazing-incidence wide-angle X-ray scattering;  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra (PDF)

## AUTHOR INFORMATION

### Corresponding Author

Karl Börjesson – Department of Chemistry and Molecular Biology, University of Gothenburg, 412 96 Gothenburg, Sweden; [orcid.org/0000-0001-8533-201X](https://orcid.org/0000-0001-8533-201X); Email: [karl.borjesson@gu.se](mailto:karl.borjesson@gu.se)

### Authors

Clara Schäfer – Department of Chemistry and Molecular Biology, University of Gothenburg, 412 96 Gothenburg, Sweden

Sandra Hultmark – Department of Chemistry and Chemical Engineering, Chalmers University of Technology, 412 96 Gothenburg, Sweden

Yizhou Yang – Department of Chemistry and Molecular Biology, University of Gothenburg, 412 96 Gothenburg, Sweden

Christian Müller – Department of Chemistry and Chemical Engineering, Chalmers University of Technology, 412 96 Gothenburg, Sweden; [orcid.org/0000-0001-7859-7909](https://orcid.org/0000-0001-7859-7909)

Complete contact information is available at:

<https://pubs.acs.org/10.1021/acs.chemmater.2c02761>

### Author Contributions

K.B. designed the project idea. C.S., together with K.B., designed the molecules. C.S. synthesized and analyzed the molecules, prepared the films, and performed spectroscopic experiments and analysis thereof. S.H. performed DSC and FSC measurements and analyzed the resulting data together with C.M. Y.Y. measured GIWAXS. All authors contributed to writing the manuscript. All authors have given approval to the final version of the manuscript.

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### Notes

The authors declare no competing financial interest.

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